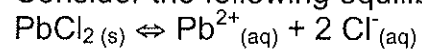


Common Ion Effect

Consider the following equilibrium system.



Describe what happens to the solubility of PbCl_2 after each of the changes.

$\text{PbCl}_2(\text{s})$ is added

$\text{Pb}(\text{NO}_3)_2$ is added

NaCl is added

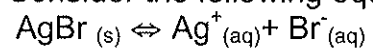
H_2O is added

AgNO_3 is added

NaBr is added

(over)

Consider the following equilibrium system.



Describe what happens to the solubility of AgBr after each of the changes are made.

AgBr(s) is added

Pb(NO₃)₂ is added

NaCl is added

H₂O is added

AgNO₃ is added

NaBr is added