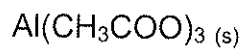
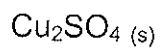
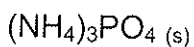


### General Ksp and Solubility → Ksp Problems

Write the equilibrium equation and solubility product Ksp for each salt. The first one is done.



The solubility of  $\text{SrF}_2$  is  $2.83 \times 10^{-5}$  M. Calculate the Ksp.

Calculate the Ksp for  $\text{CaCl}_2$  if 200.0g of  $\text{CaCl}_2$  is required to saturate 100.0 mL of solution.

Calculate the  $K_{sp}$  for  $AlCl_3$  if 100.0g is required to saturate 150.0 mL of a solution.

The solubility of  $GaBr_3$  is 15.8 g per 100.0 mL. Calculate the  $K_{sp}$ .

The solubility of  $Ag_2SO_4$  is  $1.33 \times 10^{-7}$  g per 100.0 mL. Calculate the  $K_{sp}$ .

If  $2.9 \times 10^{-3}$  g  $\text{Ca}(\text{OH})_2$  is needed to saturate 250.0 mL of solution, what is the  $K_{sp}$ ?

Calculate the  $K_{sp}$  of  $\text{CuSO}_4 \cdot \text{H}_2\text{O}$  if 63.5g is required to saturate 100.0 mL of solution.

$\text{Co}(\text{OH})_2$  Solubility =  $3.0 \times 10^{-3}$  g/L  $K_{sp}=?$

$\text{Ag}_2\text{C}_2\text{O}_4$  Solubility =  $8.3 \times 10^{-4}$  M  $K_{sp}=?$